

Skill Practice 11

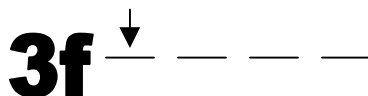
Electron Practice

Name: _____

Date: _____

Hour: _____

1. What is wrong with the following notation?



There is no such thing as an f sublevel in the third energy level. Also, an f sublevel if it existed would have five orbitals instead of four.

2. How many sublevels would you expect in the 8th energy level?

8

3. What is the maximum number of electrons that can fit in the 3d sublevel?

10

4. How many electrons can fit in a 2p orbital?

2

5. In the 5th energy level, there is a fifth sublevel called the “g sublevel”. Considering the trend in number of orbitals and electrons in the s, p, d, and f sublevels, predict how many orbitals and how many electrons can fit in a g sublevel.

Orbitals = 9

Electrons = 18

6. Considering your answer to question 5, how many electrons can fit in the entire 5th energy level?

50

7. Write the notation for an electron spinning clockwise in a p sublevel in the 4th energy level.

