

Types of Chemical Reaction Worksheet

Balance the reactions **1 to 6** and indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented:

1. $_ \text{NaBr} + _ \text{Ca(OH)}_2 \rightarrow _ \text{CaBr}_2 + _ \text{NaOH}$ Reaction Type : $_ _$
2. $_ \text{NH}_3 + _ \text{H}_2\text{SO}_4 \rightarrow _ (\text{NH}_4)_2\text{SO}_4$ Reaction Type : $_ _$
3. $_ \text{C}_5\text{H}_9\text{O} + _ \text{O}_2 \rightarrow _ \text{CO}_2 + _ \text{H}_2\text{O}$ Reaction Type : $_ _$
4. $_ \text{Pb} + _ _ \text{H}_3\text{PO}_4 \rightarrow _ _ \text{H}_2 + _ _ \text{Pb}_3(\text{PO}_4)_2$ Reaction Type : $_ _$
5. $_ _ \text{Li}_3\text{N} + _ _ \text{NH}_4\text{NO}_3 \rightarrow _ _ \text{LiNO}_3 + _ _ (\text{NH}_4)_3\text{N}$ Reaction Type : $_ _$
6. $_ _ \text{HBr} + _ _ \text{Al(OH)}_3 \rightarrow _ _ \text{H}_2\text{O} + _ _ \text{AlBr}_3$ Reaction Type : $_ _$
7. $_ _ \text{Na}_3\text{PO}_4 + _ _ \text{KOH} \rightarrow _ _ \text{NaOH} + _ _ \text{K}_3\text{PO}_4$ Reaction Type $_ _ _ _$
8. $_ _ \text{MgCl}_2 + _ _ \text{Li}_2\text{CO}_3 \rightarrow _ _ \text{MgCO}_3 + _ _ \text{LiCl}$ Reaction Type $_ _ _ _$
9. $_ _ \text{C}_8\text{H}_{16} + _ _ \text{O}_2 \rightarrow _ _ \text{CO}_2 + _ _ \text{H}_2\text{O}$ Reaction Type $_ _ _ _$

Indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented in 7 to 20.

10. $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ Reaction Type $_ _ _ _ _ _$
11. $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ Reaction Type $_ _ _ _ _ _$
12. $\text{P}_4 + 3 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_3$ Reaction Type $_ _ _ _ _ _ _ _$
13. $2 \text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be(NO}_3)_2 + 2 \text{RbF}$ Reaction Type $_ _ _ _ _ _ _ _ _ _$
14. $2 \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu(NO}_3)_2 + 2 \text{Ag}$ Reaction Type $_ _ _ _ _ _ _ _$
15. $\text{C}_3\text{H}_6\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ Reaction Type $_ _ _ _ _ _ _ _$
16. $2 \text{C}_5\text{H}_5 + \text{Fe} \rightarrow \text{Fe(C}_5\text{H}_5)_2$ Reaction Type $_ _ _ _ _ _ _ _ _ _$
17. $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$ Reaction Type $_ _ _ _ _ _ _ _$
18. $2 \text{MgI}_2 + \text{Mn(SO}_3)_2 \rightarrow 2 \text{MgSO}_3 + \text{MnI}_4$ Reaction Type $_ _ _ _ _ _ _ _ _ _ _ _$
19. $\text{O}_3 \rightarrow \text{O} \cdot + \text{O}_2$ Reaction Type $_ _ _ _ _ _ _ _ _ _ _ _$
20. $2 \text{NO}_2 \rightarrow 2 \text{O}_2 + \text{N}_2$ Reaction Type $_ _ _ _ _ _ _ _ _ _ _ _$