

Date: $\qquad$
Hour:


1. Write the symbol and charges for the following ions:

Example: calcium $=\mathrm{Ca}^{2+}$
A) phosphide $=\underline{P^{3-}}$
B) magnesium $=$ $\qquad$ C) rubidium $=\ldots \underline{R b}^{+}$
D) fluoride $=$ $\qquad$ E) aluminum $=$ $\qquad$ F) sulfide = $\qquad$
2. Each of the following formulas is written incorrectly. Please rewrite them correctly.

Example: $\mathrm{Ca}_{2} \mathrm{Cl}=\mathrm{CaCl}_{2}$
A) $\mathrm{Ba}_{2} \mathrm{~S}=\_\underline{\mathrm{BaS}}$
B) $\mathrm{Rb}_{2} \mathrm{~N}=$ $\qquad$ C) $\mathrm{Li}_{2} \mathrm{Cl}=\ldots \underline{\mathrm{LiCl}}$
D) $\mathrm{Al}_{3} \mathrm{~N}_{3}=$ $\qquad$ E) $\mathrm{Mg}_{3} \mathrm{Br}_{2}=$ $\qquad$ F) $\mathrm{O}_{3} \mathrm{Al}_{2}=\underline{\mathrm{Al}_{2}} \underline{\mathrm{O}}_{3}$
$\qquad$
3. Write the formulas for the compound formed by the combination of the following pairs of atoms.

Example: nitrogen and magnesium $=\mathrm{Mg}_{3} \mathrm{~N}_{2}$
A) sodium and sulfur $=$ $\qquad$ B) aluminum and chlorine $=\ldots \mathrm{AlCl}_{3}$
$\qquad$
C) phosphorus and calcium $=$ $\qquad$ D) barium and oxygen $=$ $\qquad$
4. For each of the formulas you wrote in question 3 above, give the names for the compounds.

Example: $\mathrm{Mg}_{3} \mathrm{~N}_{2}=$ magnesium nitride
A) $\qquad$ B) $\qquad$
C) $\qquad$
D)
$\qquad$
5. In the blanks above each column, write the charge of the ions formed by atoms in that column. Note: group IA is done for you.


