## Skill Practice 19

Name: Date:

Hour: \_\_

1. Write the symbol and charges for the following ions:

Example: calcium =  $Ca^{2+}$ 

- A) phosphide =  $\underline{P^{3-}}$  B) magnesium =  $\underline{Mg^{2+}}$  C) rubidium =  $\underline{Rb^{+}}$

- D) fluoride =  $_{\mathbf{F}}^{\mathbf{F}}$  E) aluminum =  $_{\mathbf{Al}^{3+}}^{\mathbf{F}}$  F) sulfide =  $_{\mathbf{S}^{2-}}^{\mathbf{F}}$
- 2. Each of the following formulas is written incorrectly. Please rewrite them correctly. Example:  $Ca_2Cl = CaCl_2$ 

  - A)  $Ba_2S = BaS$  B)  $Rb_2N = Rb_3N$  C)  $Li_2Cl = LiCl$

- D)  $Al_3N_3 = AlN$  E)  $Mg_3Br_2 = MgBr_2$  F)  $O_3Al_2 = Al_2O_3$
- 3. Write the formulas for the compound formed by the combination of the following pairs of atoms. Example: nitrogen and magnesium =  $Mg_3N_2$ 
  - A) sodium and sulfur =  $Na_2S$
- B) aluminum and chlorine = \_\_AlCl<sub>3</sub>\_\_\_\_
- C) phosphorus and calcium =  $\underline{\text{Ca}_3\text{P}_2}$ 
  - D) barium and oxygen = \_\_\_BaO\_\_\_\_
- 4. For each of the formulas you wrote in question 3 above, give the names for the compounds. Example:  $Mg_3N_2$  = magnesium nitride
  - A) sodium sulfide
- B) aluminum chloride
- C) <u>calcium phosphide</u>
- D) <u>bari</u>um oxide
- 5. In the blanks above each column, write the charge of the ions formed by atoms in that column. Note: group IA is done for you.

